

1. Complete the following chart.

	Where located in atom	electrical charge	mass in AMU
Proton	_____	_____	_____
Electron	_____	_____	_____
Neutron	_____	_____	_____

2. Calculate the number of protons, electrons and neutrons there are in the following elements. Use a periodic table to help.

<u>substance</u>	<u>protons</u>	<u>electrons</u>	<u>neutrons</u>	<u>mass #</u>
Calcium (Ca)	_____	_____	_____	<u>40</u>
Iron (Fe)	_____	_____	_____	<u>55</u>

Look below at the number of subatomic particles in atom #1 and atom #2:

atom #1 has: 15 protons 15 electrons 16 neutrons

atom #2 has: 17 protons 17 electrons 18 neutrons

3. Use your periodic table to identify the element each atom represents.

_____atom #1 _____atom #2

4. Explain how you decided which element you chose for question #3.

5. What is the mass number of each element?

_____atom #1 _____atom #2

6. Draw each atom in #3-5 showing the neutrons and protons in a nucleus and how the electrons would look in their orbits.